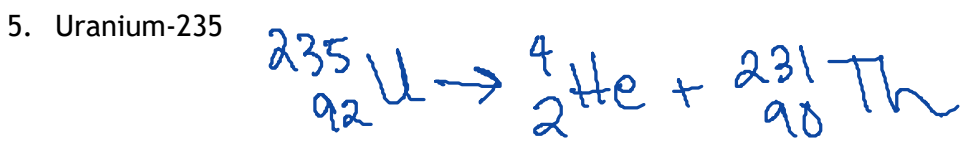
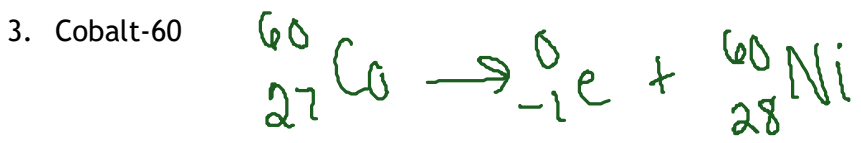
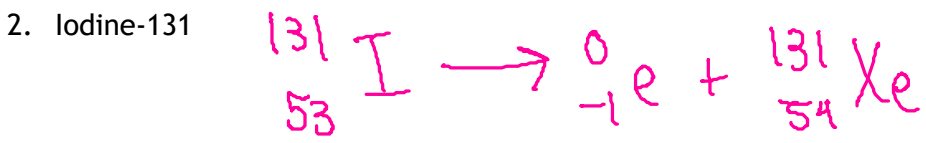


Name: _____ Date: _____

Balancing Nuclear Decay Equations Practice

Please write each of the following decay equations:

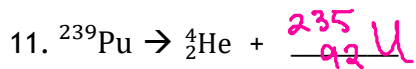


Please complete each of the following nuclear decay equations by identifying the missing component and indicate if the equation is artificial or natural transmutation.

<u>Nuclear Equation</u>	<u>Natural or Artificial Transmutation</u>
6. ${}_{94}^{240}\text{Pu} \rightarrow {}_{92}^{236}\text{U} + {}_2^4\text{He}$	<u>Natural Transmutation</u>
7. ${}_{13}^{27}\text{Al} + {}_{+1}^0\text{e} \rightarrow {}_{14}^{27}\text{Si}$	<u>Artificial Transmutation</u>
8. ${}_{84}^{210}\text{Po} \rightarrow {}_{+1}^0\text{e} + {}_{85}^{210}\text{At}$	<u>Natural Transmutation</u>
9. ${}_1^2\text{H} + {}_1^2\text{H} \rightarrow {}_2^4\text{He}$	<u>Artificial Transmutation</u>
10. ${}_{83}^{209}\text{Bi} \rightarrow {}_{82}^{208}\text{Pb} + {}_{+1}^0\text{e}$	<u>Natural Transmutation</u>

Name: _____ Date: _____

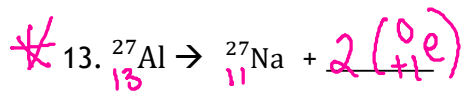
Balancing Nuclear Decay Equations Practice



Natural Transmutation

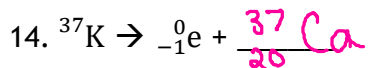


Natural Transmutation

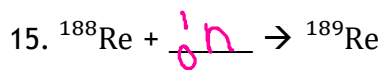


2 positrons released

Natural Transmutation



Natural Transmutation



Artificial Transmutation